

## Selective Skeletal Rearrangement by Carbon–Carbon Bond Activation

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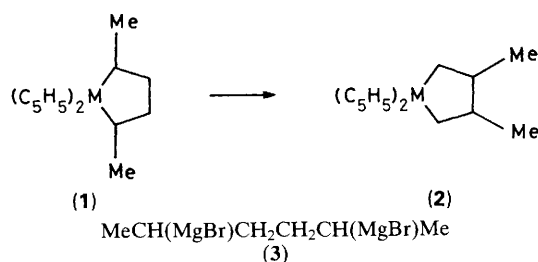
Skeletal rearrangement of  $\alpha$ -substituted organometallic compounds to give  $\beta$ -substituted compounds by carbon–carbon bond activation was performed with bis( $\eta^5$ -cyclopentadienyl)-2,5-dimethyl-zircona- or -hafna-cyclopentane, which were regio- and stereo-selectively converted into the corresponding 3,4-dimethylmetallacyclopentanes in high yield.

Selective skeletal rearrangement by carbon–carbon bond activation is a potentially attractive methodology in organic synthesis. Reaction (1) for organometallic compounds can be considered to be representative. We now describe a novel selective rearrangement by carbon–carbon bond activation using organozirconium or organohafnium compounds (Scheme 1).

The zirconium complex (**1a**) with methyl groups on the  $\alpha$ - and  $\alpha'$ -carbon atoms was prepared *in situ* from  $(C_5H_5)_2ZrCl_2$  and the corresponding Grignard reagent (**3**), and underwent selective skeletal rearrangement during 1 h at room temperature by carbon–carbon activation to give 2,3-dimethylbutane after hydrolysis in 98% yield based on the metal. Bromination of the reaction mixture instead of hydrolysis gave ( $\pm$ )-1,4-dibromo-2,3-dimethylbutane selectively [( $\pm$ ):*meso* 99:1] (Scheme 2). 2,5-Dibromohexane was totally absent.

The hafnium derivative (**1b**) reacted similarly, at higher temperature. The rearrangement product 2,3-dimethylbutane was obtained in 53 and 98% yield during 1 h at 50 and 80 °C,

respectively, after hydrolysis. The reaction was monitored by  $^1H$  NMR spectroscopy. Reaction of  $(C_5H_5)_2HfCl_2$  with the Grignard reagent (**3**) quantitatively afforded (**1b**), the  $^1H$  NMR spectrum of which showed a  $C_5H_5$  signal at  $\delta$  5.89. At 50 °C three new singlets assignable to the  $C_5H_5$  groups of (**2b**) and two intermediates, 2,4-dimethylhafnacyclopentane isomers, emerged at  $\delta$  5.93, 5.90, and 5.87, respectively, and grew at the expense of that of (**1b**) at  $\delta$  5.89. This process obeyed first-order kinetics,  $k_1 = (6.5 \pm 0.2) \times 10^{-3} \text{ min}^{-1}$  at 50 °C. The final spectrum revealed only one singlet peak for



a: M = Zr  
b: M = Hf

Scheme 1

